

Introduction to Chemistry  
Quiz 4 c

Name: \_\_\_\_\_  
Date: \_\_\_\_\_

Atomic Structure  
Electron Configuration  
Periodic Trends  
Ions  
Ionic and Covalent Bonds

1. Indicate the number of valence electrons for an atom of
  - a) Selenium \_\_\_\_\_
  - b) Gallium \_\_\_\_\_
2. Provide the predicted charge for the most common ions of
  - a) Gallium \_\_\_\_\_
  - b) Selenium \_\_\_\_\_
- 3) Which of the following are most likely to form negatively charged ions?
  - a) Metals
  - b) Metalloids
  - c) Nonmetals
  - d) Noble gases
  - e) None of the above
- 4) Which element would have the largest ionization energy?
  - a) Ni
  - b) N
  - e) Ne
  - c) Na
  - d) H
- 5) Write the full or the shorthand electron configuration for the element phosphorous.

Fold quiz with your work inside, then write your name on the outside.